

Pitting Growth Rate in Carbon Steel Exposed to Simulated Radioactive Waste (U)

by

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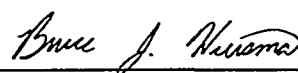
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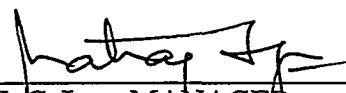
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PITTING GROWTH RATE IN CARBON STEEL EXPOSED TO SIMULATED RADIOACTIVE WASTE

SUMMARY

Dilute high-level radioactive waste slurries can induce pitting corrosion in carbon steel tanks in which such waste is stored and processed. The waste is normally maintained with closely monitored nitrite and hydroxide concentrations known to prevent the initiation of pitting. Coupon immersion tests are being conducted in laboratory simulants of waste to determine the probability and growth rate of pitting in steel in the event of out-of-limits nitrite concentrations. Sets of about 36 carbon steel coupons have been immersed in known corrosive conditions (nitrite < 5% of the established limit) at a temperature of 50°C. Three sets have been removed from testing after 64, 150, and 350 days of immersion. The long immersion times introduced variability in the exposure conditions due to the evaporation and replenishment of solution. The deepest corrosive attack was measured on each coupon by optical microscopy. The deepest pits were ranked and analyzed as a type 1 extreme value distribution to extrapolate from the coupon population to the maximum expected pit depths in a waste tank structure. The data were compared to a power law for pit growth, although the deepest pits did not increase monotonically with time in the limited data set.

INTRODUCTION

Localized corrosion is a broad term that describes the forms of corrosion in which the attack is not evenly distributed over a surface. Localized corrosion has been defined to include such distinct forms as pitting, crevice corrosion, intergranular corrosion, stress corrosion cracking, corrosion fatigue, cavitation attack, and others.¹ Some investigators have defined the term more narrowly. Uhlig identifies localized corrosion as either pitting or crevice corrosion,² and this definition generally prevailed at a recent conference on advances in the understanding of localized corrosion.³ The industrial impact of localized corrosion is significant. Based on the inclusive definition, an analysis of metallic failures at one American chemical plant revealed that localized corrosion mechanisms were responsible for about two thirds of the failures, while general corrosion was responsible for fewer than one fourth.⁴ A large fraction of the localized corrosion failures occurred by pitting.

The definition of pitting rests upon the concept of passivity in metals. Pitting is the local electrochemical dissolution of metal from a surface that is electrochemically passive. Passivity arises from the growth of a stable, non-porous oxide film that markedly impedes the rate of general dissolution of the alloy in reactive environments. Pitting is inherent in the performance of engineering alloys, such as iron- and aluminum-based alloys, because these alloys derive their utility and stability in potentially corrosive environments from the presence of a passive oxide film on their exposed surfaces. Pitting may begin with a small, randomly located breach of the passive film, which exposes the underlying alloy to the environment and active corrosion. The anodic reactions in the initiated pit can occur at a

very high rate because the supporting cathodic reactions take place on the large remaining passive surface. Most of the surface undergoing pitting continues to be passive. The ratio of pitting current density to passive surface current density may be as high as one million.⁵

Pitting is perhaps the most difficult corrosion form to evaluate and characterize in an engineering structure or component, because of the randomness of its initiation and propagation. Ives has defined five stages in the development of a corrosion pit: (1) local breakdown of the passive film, (2) high anodic dissolution current density, (3) competition between repassivation and establishment of continuing pit growth, (4) growth from atomic scale to microscopic scale, and (5) growth from microscopic to macroscopic size, or simply, macroscopic pitting.⁶ The literature on pitting is extensive and largely focussed on its early stages.⁷ From the point of view of engineering service of materials, macroscopic pitting is of greatest interest. It is typically the through-thickness macroscopic pit that results in the failure of a structure. However, stage (5) pit growth has been studied far less than the earlier stages. This is probably due to the random nature of the growth, growth rate, and cessation of growth and the cost and technical challenges in the study of macroscopic pitting. As will be discussed below, the successful approach to understanding macroscopic pit growth is based on the statistical analysis of measurements of large populations of pits.

Prevention of Pitting in High-Level Radioactive Waste Solutions

High-level radioactive waste at the Savannah River Site (SRS) is stored in large carbon steel tanks. The process for vitrifying the waste requires that the waste slurries be washed to remove soluble salts. Washing will be conducted in existing carbon steel tanks similar to those in which the waste is stored. Washing lowers the concentration of hydroxide, which in stored waste is maintained at a concentration sufficient to prevent pitting and stress corrosion cracking. Near the liquid waste/vapor interface and in an aqueous film on the tank wall above the interface, the hydroxide concentration is decreased below its bulk value by reaction with absorbed atmospheric carbon dioxide. This reaction proceeds rapidly, and the diffusion of hydroxide from the bulk liquid to the film is slow enough to result in a steady-state hydroxide concentration in the film that is 3 to 4 orders of magnitude lower than the bulk concentration. The tank wall and the tank cooling coils at the liquid-vapor interface are the areas most vulnerable to pitting. Hydroxide is also destroyed by radiolysis and consumed through reaction with organic acids in certain waste streams. Over a few years' time, the hydroxide concentration of the entire volume of washed waste slurry would decline to the low steady-value in the absence of replenishment, and the waste would become corrosive to the steel. Thus, through dilution and chemical reaction, the hydroxide concentration in washed waste slurries is insufficient to prevent pitting corrosion.

The technical approach to the control of pitting corrosion has been to modify the waste chemistry to prevent the initiation of pitting. Prevention of initiation was deemed a more tractable technical problem than slowing pit propagation to a safe rate, due to the difficulty of obtaining reliable rate data. Extensive laboratory studies, using electrochemical and coupon immersion test methods, have been conducted to identify a corrosion inhibitor that will supplement the remaining hydroxide to prevent pitting. Through these studies, nitrite

was shown to be an effective inhibitor and the concentrations of nitrite necessary to prevent the initiation of pitting in dilute waste slurries were determined.^{8,9} Various nitrite levels were tested throughout the range of possible waste compositions and temperatures. Nitrite was shown to depend upon the temperature and a single aggressive anion concentration. The principal aggressive anion is generally nitrate. Where the nitrate concentration is reduced by radiolysis to a sufficiently low value, sulfate or chloride may induce pitting in carbon steel and control the required nitrite level.⁸ Equations relating the safe nitrite level to the waste temperature and its nitrate, sulfate, or chloride concentration are incorporated into the process requirements for the vitrification feed preparation tanks.

The process requirement nitrite equations are conservative. They incorporate a 50% safety factor over the equations fitted to the experimental data, and the experimental data were conservatively measured. This conservatism has been scrutinized because of undesirable downstream effects of nitrite in the vitrification feeds. These effects have resulted in the construction of a new stainless steel facility in which nitrite can be washed out of one of the feeds. The present study was undertaken partly in response to the need to understand the extent of the potential pitting as it impacted the conservative nitrite inhibitor concentrations. The approach of the present work has been to determine the pitting response in an intentionally corrosive dilute waste solution and hence to determine the probability and rate of pitting so as to establish the technical basis for out-of-limits operating conditions.

Macroscopic Pit Growth in Carbon Steel

The kinetics of pit growth are widely described with a power law of the form

$$P = kt^n \quad (1)$$

where P is pit depth, t is time, k is a constant dependent upon material and environment, and n is a numerical constant.⁷ This form was proposed to describe the pits measured on numerous ferrous materials in the very large National Bureau of Standards underground corrosion program that was conducted from 1922 to 1952. The program was summarized in the frequently cited monograph by Romanoff.¹⁰ The exponent n was usually in the range 0.5 to 0.6, but varied widely from near zero to 0.92. The constant k varied from 20 to 107. The constants are empirical quantities dependent upon the alloy and environment in test. Equation 1 is not predictive in a quantitative sense in the NBS underground corrosion survey, which is perhaps expected given the wide range of environments in which the data were acquired. Nonetheless a review of the NBS data along with other long-term corrosion databases concludes that Equation 1 fits the data as well or better than other relationships.¹¹ In other applications of Equation 1 the value of n was determined to be 0.33 in aluminum and 0.5 in mild steel.⁶ It varied from 0.3 to 0.7 in alloy 600 in a temperature dependent manner.¹² More recent work on a mild steel developed the following area-adjusted equation for the maximum pit depth on a radioactive waste canister

$$P = 8.35t^{0.46} \quad (2)$$

where depth P is in millimeters and t is the time in years. The steel was immersed in slightly alkaline (pH 8.4) synthetic granitic ground water at 90 °C, and the corrosive agent was chloride.¹³ The growth of pits in a low alloy steel exposed to deionized water at 65°C resulted in a power law of

$$P = 13.65t^{0.584} \quad (3)$$

where P is in micrometers and t in hours.¹⁴

The NBS studies investigated the relationship between the size of the specimen or structure under study with the depth of deepest pits or corrosive attack. It was concluded from the NBS observations that the depth of the deepest pits P_d increases with increasing specimen area A according to

$$P_d = bA^m \quad (4)$$

or

$$P_d = c + d \ln A \quad (5)$$

where b , c , d , and m are constants. Equations 4 and 5 were developed from the observations that the larger the specimen, the deeper is the pit likely to be measured.

Extreme Value Theory

As stated above, in many cases the few deepest pits are of much greater interest than the many shallower pits. If pitting is random and there is a statistical distribution of pit depths, then the deepest pits themselves conform to a statistical distribution. It usually is not possible to determine the underlying distribution of all pit depths because the smallest cannot be readily measured. However the statistics of extreme values have been found useful in analyzing and characterizing the distribution of deepest pits. The development of extreme value theory for engineering applications is credited throughout the corrosion literature to Gumbel.¹⁵ In particular, extreme value theory has provided the mathematics to predict the deepest corrosive attack in a large structure from a limited number of measurements on a portion of the structure, or alternatively, the deepest corrosive attack on a large number of coupons from measurements on only a fraction of that number. Gumbel's work was first applied to pitting corrosion by Aziz, who used the theory to analyze pit depths on aluminum coupons and pipes exposed to potable and sea water¹⁶ and by Eldredge, who analyzed pits in oil well tubing.¹⁷ There were subsequent applications of extreme value theory to pitting on steel pipelines.¹⁸ Interest in the long-term performance of radioactive waste containers stimulated more recent use of extreme value theory, including studies at the National Institute of Standards and Technology (formerly the National Bureau of Standards).

McNeil has reviewed some statistical issues associated with carbon steel pitting for

radioactive waste storage applications.¹⁹ He considered especially how extreme value theory may properly be applied to pitting data. There are three types of distributions cited in the corrosion literature: type 1, exponential; type 2, Cauchy; and type 3, limited. These may be differentiated by the bounds placed on allowed measurements. In type 1 distributions, there are no bounds to the data, whereas in type 2 and type 3 distributions, the measured data are bounded by a minimum and a maximum allowed value, respectively.

Mathematically, the type 1 distribution is the simplest, having an exponentially decreasing probability of occurrence with increasing value of the variate. Many common distributions, such as the exponential and normal distributions, converge asymptotically at extreme values to the type 1 distribution. Its form is the double exponential

$$F(x) = e^{-e^{-y}} \quad (6)$$

with the reduced variate y given by

$$y = \alpha(x - u) \quad (7)$$

where y lies in the range $-\infty < y < \infty$, u is a parameter defining the location of the distribution, and α a parameter defining the scale (or width) of the distribution. $F(x)$ is interpreted as the probability that any pit measured is shallower than a depth x . The quantity $1 - F(x)$ then is the probability of measuring a pit deeper than x , or, alternatively, the risk of accepting x as the deepest possible pit. The reciprocal of $1 - F(x)$ is called the "return period" and will be discussed further below.

As stated above, the early use of extreme value theory in pitting analyses is credited to Aziz and Eldredge. They interpreted their data with the exponential type 1 distribution. The unbounded range of this distribution is somewhat unphysical in the field of corrosion because it does not reflect the finiteness of the corroding structure or the electrochemical limits to pitting that are enforced by the IR drop through an increasing long electrolyte circuit path. Nonetheless the type 1 distribution has been acknowledged to yield satisfactory results.

EXPERIMENTAL

Coupon immersion tests were conducted with ASTM A537 Class 1 carbon steel coupons in a non-radioactive simulant of a dilute sludge slurry expected in Extended Sludge Processing in Tanks 40H, 42H, and 51H. The simulant composition is shown in Table 1. Two 36-liter volumes of solution were mixed from reagent-grade chemicals and distilled water. The solutions were adjusted to a pH of 9.73 with sodium hydroxide. The total nitrate concentration was 0.0546 M, with the nitrate ion contributed mainly by sodium nitrate, aluminum nitrate, and ferric nitrate. The corrosion database indicates that, for this nitrate level, the minimum nitrite concentration required to prevent pitting is 0.16 at the planned immersion temperature of 50°C.⁸ The actual nitrite concentration in the simulant

was 0.005 M to ensure that the solution would induce pitting.

The coupons used in this study were fabricated from ASTM A537 class 1 carbon steel and supplied by Metal Samples, Inc., Munford, Ala. Their size was 6.0 in. by 2.0 in. by 0.12 in., and they were used in the as-received condition with 600 grit surface finish. The large coupons provided the surface area on which significant pitting would be measurable. Two coupons were placed in 1-L plastic bottles, which were filled with test solution to a mark to cover the coupons. The entire test consisted of 70 bottles containing 140 coupons. The bottles were placed in convection ovens maintained at $50 \pm 2^\circ\text{C}$. Sets of coupons were scheduled for periodic removal. Some evaporation occurred, leading to the exposure of some coupons to a liquid/vapor interface. The bottom of the coupons rested in settled solids (ferric hydroxide, manganese dioxide) which, along with the coupon contact with the bottle bottom, created a crevice effect. In some instances, one coupon rested against the other, which created a large crevice. As will be noted below, the interface and crevice regions on some coupons were the sites of deep corrosive attack.

Upon removal from the test, the coupons were cleaned of corrosion products in either an inhibited hydrochloric acid solution or in a sodium hydroxide solution/zinc mixture, according to ASTM standard practice. Considerable corrosion product remained and was removed mechanically from pits with either a sharp steel stylus or a carbide-tipped vibratory engraving tool. The coupons were then examined with optical microscopy, and depth of corrosion attack was measured with a calibrated focus microscope. The standard practice for pit depth measurement was first to focus the microscope on the uncorroded metal near the location of interest and then zero the digital readout of the dial micrometer attached to the microscope stage. The bottom of the pit was then brought into focus and the depth read. The sensitivity of the measurement technique was 0.1 mil (2.5 μm). However, the uncertainty in the point of sharp focus for the bottom resulted in an uncertainty in the measurement of about ± 0.3 mils (8 μm). Selected coupons were also sectioned to measure pit depths by serial metallography.

RESULTS

The present results consist of maximum pit depths from three sets of coupons that were immersed for 64, 150, and 350 days, respectively. As expected, the test solution produced significant corrosion. Most coupons had regions of what has been termed uneven local corrosion,²⁰ with otherwise unattacked metal between the regions. Few coupons had isolated small pits only. This corrosion was located generally (1) at the liquid/vapor interface near the top of the coupon, (2) down the length of the coupon in roughly vertical trailing bands, and (3) at the coupon bottom in contact with the bottle or with solids. The vertically trailing corrosion has been frequently noted as evidence of the ability of corrosion products to induce further corrosion as gravity draws them down to fresh metal surface.²⁰ The corrosion products were primarily magnetite with some goethite as revealed by x-ray diffraction.

Within the regions of uneven local corrosion, there was aggressive attack that produced pits markedly deeper than the adjacent corrosion. A wire brush, steel scribe, and vibratory

engraver were used to dislodge the rust that remained after chemical cleaning. Very little underlying uncorroded steel was disturbed. This procedure was followed on both front and rear faces of the coupon. With the rust removed, the coupons were examined with a binocular microscope at magnifications up to 68X to identify deepest pits for measurement at magnifications up to 400X. Pits that were near, and influenced by, an edge of a coupon were not included in the measurements. The results of the deepest pit measurements are shown in Table 2. The measurements ranged from 1.0 mil (25 μm) to 31.3 mil (800 μm) in the 64-day data set, from 1.0 mil (25 μm) to 49.4 mil (1.250 μm) in the 150-day set, and from 8.7 mil (220 μm) to only 34.3 mil (870 μm) in the 350-day set. The data are shown in Table 2, where the coupons are identified by number and ranked in order of increasing depth of the deepest pit.

It is interesting to contrast these measurements with those of Marsh et al., which were made on mild steel in slightly alkaline synthetic ground water at 90°C. Those measurements ranged from about 80 mils (2000 μm) to 320 mils (8000 μm) in the time period 40 days to 400 days.¹³ The electrolyte was similar to that of the present work. The greater corrosive attack is consistent with the higher test temperature.

The depths have been analyzed as a type 1 extreme value distribution. The procedure of Aziz was used to assign a plotting position on the ordinate for the relevant measured depth on the abscissa. This position is the cumulative frequency of the distribution and is obtained by dividing the rank of the measurement, m , by the sum of 1 plus the total number of coupons N . The cumulative frequency can be converted to the linear scale of the reduced variate y of Equation 5 by a double logarithmic transformation

$$y = -\ln \left[-\ln \left(\frac{m}{N+1} \right) \right] \quad (8)$$

This converted cumulative frequency has been plotted against the m th pit depth in Figure 1. A least squares fitting routine in Microsoft Excel™ software was used to fit linear equations to the data. The results of the fits to each immersion period are shown in Table 3, along with the appropriate R^2 coefficient. The latter are clustered about 0.9.

The linear equations can be usefully extrapolated to investigate the likelihood of a deepest pit on larger surface areas. This procedure uses the concept of the return period. In practical terms, the return period, T , is "the number of coupons that must on average be exposed in order to obtain a pit depth greater than the observed pit depth."¹⁴ In the present work, the coupon results are to be applied to a large carbon steel waste tank. Thus, the return period is the area of the waste tank that is vulnerable to pitting corrosion relative to the area of the coupon exposed in the experiment. In the analysis the entire area of the coupon is considered.

The area of a coupon was 24.0 in.² (0.0155 m²). The area of primary tank wall of an 85-ft (26-m) diameter waste tank (0.50-in. [12-mm] wall thickness) that is subject to corrosive

conditions of washed sludge slurry is 19200 in.^2 (12.4 m^2), based on an aqueous film extending 6 in. (0.015 m) above the liquid/vapor interface. Dividing the coupon area into the waste tank area yields a value of 800 for the return period T . The expected deepest pits are said to converge to

$$x = \mu + \frac{\ln T}{\alpha} \quad (9)$$

for type 1 extreme value distributions.¹⁴ Inserting the values for T and the distribution parameters yields the expected deepest pits for the experimental periods. These are listed at the bottom of Table 3. The deepest pit expected to develop in a tank in a 64 -day period of aggressive conditions is 36 mils ($910 \mu\text{m}$); in a 150-day period the pit is expected to be 50 mils ($1300 \mu\text{m}$) and in 350 days, 30 mils ($760 \mu\text{m}$).

The expected maximum pit depths may be analyzed to develop an expression for the time dependence of pit growth. There are only three values at this stage of the work, and it will be noted that the expected maximum pit depths do not increase monotonically with time. The low maximum pit depth for the 350-day set of data may be due to the variation in exposure conditions caused by evaporation and replenishment of test solution over that long test period. Although the time dependence is expected to follow Equation 1, Figure 2 shows that such a relationship does not provide a good fit to the present data. The curve in Figure 2 was obtained by an iterative calculation of the minimum square of the residuals, with constants k equal to 3.6 and n equal to 0.45. With these constants the expected deepest pit after nearly a year is 50 mils, which represents only 10% of the tank wall thickness and 33% of a tank cooling coil thickness. This power law fit will be reanalyzed when pit depth measurements are made after immersions of up to 730 days' duration.

CONCLUSIONS

A type 1 extreme value distribution has been used to analyze deepest pits measured in a large coupon immersion experiment on carbon steel in simulated alkaline radioactive waste. Good correlations with the type 1 distribution are obtained on individual sets of deepest pit data from immersion periods of 64, 150, and 350 days. These distributions have been used to extrapolate the deepest pits measured in the population of corrosion coupons to an expected maximum value on the primary wall of a waste tank. However, the maximum depths measured do not increase monotonically with time. The widely used power law for pit growth was applied to the present limited data set, but adherence to the power law was not demonstrated.

Pitting in carbon steel tanks containing dilute alkaline radioactive waste is prevented at SRS by closely monitored additions of sodium nitrite. The present work indicates that, in simulated wastes containing less than 5% of the minimum nitrite concentration for an ESP solution at 50°C and 0.0546 M nitrate, the pitting growth rate in carbon steel is low enough to present no risk of penetration of a waste tank wall within one year. The maximum depth after one year under these conditions is predicted to be only 10% of the one-half-inch tank wall thickness. A more reasonable out-of-limits nitrite concentration of, for example, 80%

of the process requirement limit should be even less aggressive and result in a smaller maximum pit depth than 50 mils.

FUTURE WORK

The results reported here indicate that, in the tested time period with an extremely low nitrite concentration, pitting of the order of ten mils might develop, before the restoration of inhibitor levels. However, the response of a population of growing pits to the restoration of required nitrite levels is not known. The considerable corrosion product in and above a pit may prevent the migration of inhibitor ions to the active corrosion site at the bottom of the pit. Thus, the pit may continue to grow even as the bulk environment has acquired effective inhibitor concentrations.

This possible lack of communication is the basis for the assumption of a five-day limit as the basis for calculating the amount of uninhibited water that may be added to a waste tank.²¹ It was judged that pits forming within five days would be shallow enough to be repassivated by the inhibitors diffusing into the specified maximum uninhibited water volume from the bulk environment. This judgment rests upon the experience of cooling coil failures during waste removal in the 1960s.²²

Additional coupon immersion tests are planned to address the effect of the restoration of an effective nitrite concentration on a population of active pits. The results of these tests will help establish the technical basis for the time to recover from an out-of-limits condition. The additional tests will be started after the pit growth rate has been accurately determined.

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TABLE 1
COMPOSITION OF SLUDGE SIMULANT

Compound	Concentration (M)
Sodium Carbonate	1.73E-02
Sodium Bicarbonate	4.46E-02
Sodium Nitrite	5.00E-03
Manganese Dioxide	5.75E-03
Aluminum Nitrate	1.50E-03
Sodium Oxalate	1.94E-04
Sodium Chromate Anhydrous	7.94E-05
Sodium Molybdate	1.03E-05
Sodium metasilicate	7.86E-05
Sodium Phosphate Tribasic	2.23E-04
Chromic Chloride	3.75E-05
Cupric Nitrate	4.30E-05
Ferric Nitrate	1.25E-03
Nickel Nitrate	1.50E-03
Mercuric Nitrate	2.50E-04
Cobalt Nitrate	3.00E-05
Sodium Chloride	7.14E-04
Sodium Fluoride	4.13E-04
Sodium Sulfate	3.63E-03
Sodium Nitrate	4.27E-02

TABLE 2

Rank	Data from 64-Day Immersion		Data from 150-Day Immersion		Data from 350-day immersion	
	Coupon No.	Deepest Pit	Coupon No.	Deepest Pit	Coupon No.	Deepest Pit
1	60	1.0	81	1.0	33	8.7
2	52	1.7	86	7.0	47	9.5
3	21	2.0	76	11.0	23	11.3
4	2	3.2	82	11.7	25	11.5
5	7	3.5	73	12.4	46	12.4
6	61	4.1	104	12.4	70	13.7
7	54	6.1	87	13.2	59	14.4
8	4	6.3	88	13.3	48	14.8
9	6	7.6	83	13.4	29	14.8
10	53	8.6	93	13.4	28	15.0
11	35	8.7	102	14.2	30	15.6
12	14	10.2	105	16.2	49	15.6
13	64	10.6	75	17.0	42	16.3
14	63	10.7	91	17.3	24	16.9
15	8	11.9	74	17.4	44	17.0
16	20	12.1	79	18.2	58	17.2
17	51	12.3	94	18.3	1	17.4
18	68	12.5	85	18.5	11	18.3
19	15	12.6	92	18.5	41	18.5
20	69	13.0	101	18.8	57	18.9
21	17	13.1	78	19.8	50	19.2
22	3	13.2	77	20.7	39	19.5
23	13	13.2	103	21.1	43	19.8
24	37	13.3	106	21.8	31	20.1
25	27	13.4	99	21.9	56	20.4
26	36	13.6	71	22.3	45	20.6
27	9	13.7	89	22.3	52	21.9
28	12	14.1	96	22.6	55	23.0
29	34	14.6	90	22.9	40	23.2
30	62	15.9	100	23.2	18	24.5
31	26	16.2	84	23.3	19	25.3
32	16	18.3	72	25.0	10	31.4
33	5	19.5	80	25.9	22	34.3
34	67	20.1	98	26.0		
35	66	22.4	97	40.8		
36	65	31.3	95	49.4		

TABLE 3
RESULTS OF LEAST SQUARES FITS OF
TYPE 1 EXTREME VALUE DISTRIBUTION

<u>Parameter</u>	<u>64-Day Data</u>	<u>150-Day Data</u>	<u>350-Day Data</u>
Slope, α	0.18	0.13	0.20
Intercept, μ	-1.56	-1.92	-3.15
R ² Coefficient	0.94	0.88	0.88
Expected Deepest Pit in mils (μm)	36 (910)	50 (1300)	30 (760)

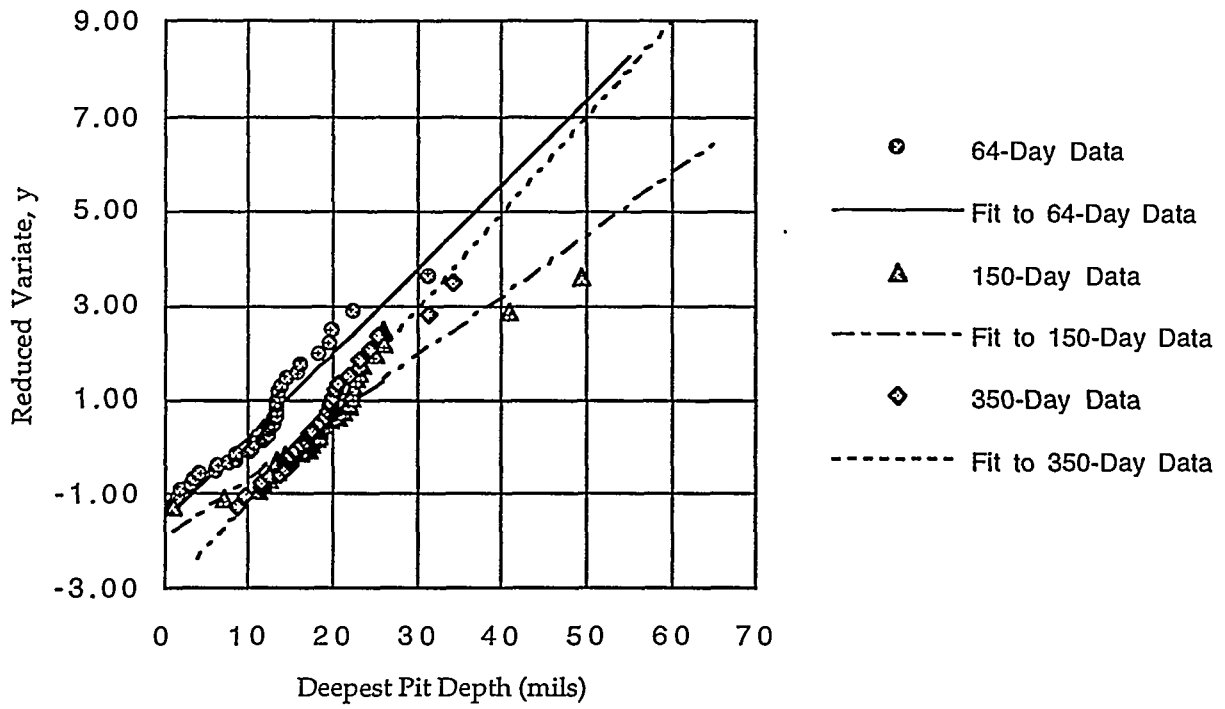


Figure 1. Type 1 extreme value distribution plot of the deepest pit data.

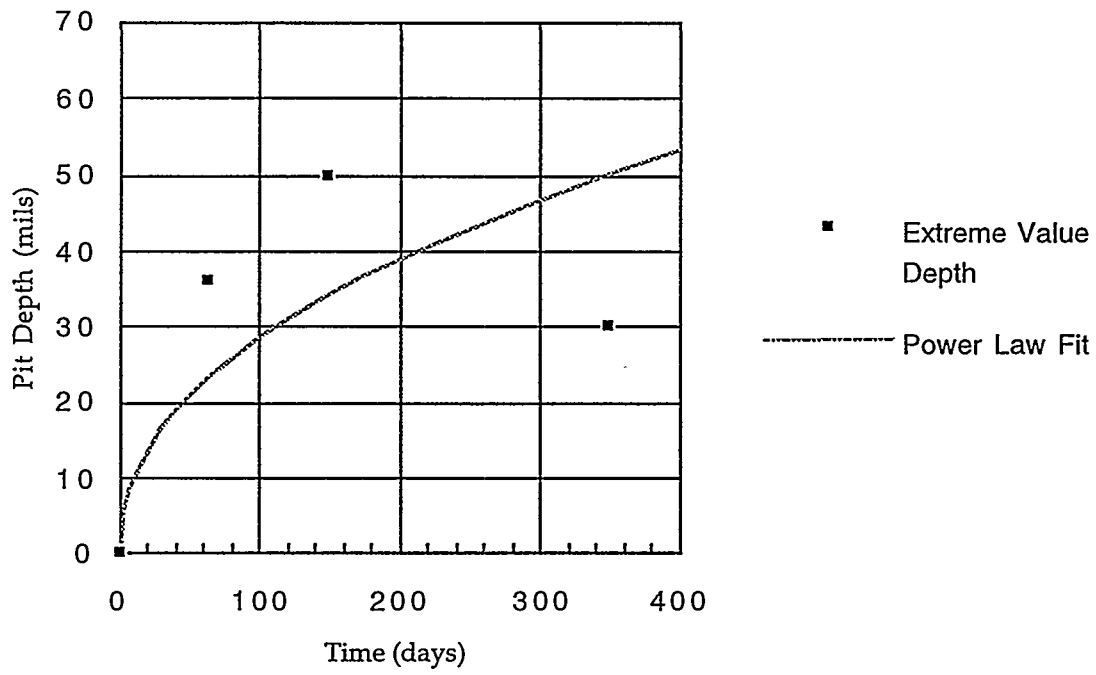


Figure 2. Expected deepest pits and a typical power law fit for carbon steel pitting.

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Distribution List

G. T. Wright, 703-H
T. J. Lex, 719-4A
H. D. Harmon, 719-4A
T. M. Monahan, 703-H
V. G. Dickert, 703-H
B. L. Lewis, 703-H
J. E. Marra, 241-120H
E. D. Lee, 241-152H
T. C. Hsu, 703-H
J. A. Pike, 703-H
J. D. Cohen, 730-A
T. L. Capeletti, 773-41A
N. C. Iyer, 773-A
L. M. Papouchado, 773-A
W. L. Tamosaitis, 773-A
S. D. Fink, 773-A
D. T. Hobbs, 773-A
C. F. Jenkins, 730-A
J. I. Mickalonis, 773-A
B. J. Wiersma, 773-A
MTS Records, c/o V. W. Anderson, 773-41A
Technical Information Management, 703-43A (4)
High Level Waste File, 703-H [File Code: WM 1010.0]